There are three types of intermolecular forces: London dispersion forces, dipole–dipole interactions and hydrogen bonding Intermolecular forces can be used to predict relative boiling points.H>dipole bonding>London dispersion (van der Waals) 3:06 a.m Dipole forces of polar molecules generally result from the attraction of the positive and negative ends of the dipole moments of polar molecules such as CH3OCH3 London dispersion forces that affect all particles, and London dispersion forces between them Nonpolar molecules generally arise from the induced temporary dipole moments in molecule by other nearby molecules such as ethane C2H6 Hydrogen bonding is not a true bond but is a particularly strong dipole attraction.CH3CH2OH is stronger than CH3CH2NH2