

Addition of strong acid or base to buffers When strong acid is added to a buffer solution containing HA and A^- , the H^+ consumes A^- to form HA. So, $[HA]$ increases while $[A^-]$ decreases. Two steps are required to calculate the resulting pH – first, a stoichiometry problem for the neutralization; second, an equilibrium problem for the new concentrations. Here, $[HA]$ decreases while $[A^-]$ increases.