Similarly the concentration of hydroxide ion, OH−, of a solution is commonly expressed in term of the pOH of the solution , which is define as the negative logarithm of OH− pOH = −log (OH−) The hydroxide ion concentration can be obtained from pOH of the solution using the equation : (OH−) = 10− pOH Additionally the pH and pOH of any aqueous solution are related as are the hydrogen and the hydroxide ion concentration , the relative equation are : (H+)(OH−) = 1×10−14 pH + pOH = 14