Introduction In this practical work, we focus on the determination of the unknown concentration of oxalic acid (HECEOE) through a redox titration process. The principle underlying this experiment hinges on a redox reaction, where transfer of electrons occurs between oxidizing and reducing agents. Specifically, oxidation involves electron loss, while reduction pertains to electron gain. By employing potassium permanganate (KMnOE) as the titrant, and using the titration method without an external indicator due to the self-indicating nature of KMnOE, we aim to identify the precise point of reaction through a color change. This practical exercise is crucial for understanding redox processes and their applications in .analytical chemistry