In coordination chemistry, tetrahedral complexes are usually high spin due to several intrinsic factors related to their geometric and electronic structure: 1.**Electronic Configuration**: The electrons prefer to occupy higher energy orbitals rather than pair up in lower energy orbitals because pairing up requires overcoming the electron-electron repulsion, which is not sufficiently compensated by the weak crystal field splitting.These factors contribute to tetrahedral complexes generally exhibiting high spin states.This results in a weaker crystal field splitting.Typically, ?t is about 4/9 of ?o.2.3.