

The mole is a unit of measurement used to describe the amount of a chemical species. One mole contains 6.022×10^{23} particles (atoms, molecules, ions, electrons). $n = \frac{\text{mass of substance (g)}}{\text{molecular mass}} = \frac{m}{MW}$ Molarity: In chemistry, molarity is the most frequently used method of expressing concentration of a solution. It can be used to describe the number of atoms, molecules, ions, electrons, etc. To determine the number of moles, n , in a given quantity of a substance, you divide the given quantity of the substance by the molecular weight.