As the hydrolysis proceeds, there will be proportional proport increase in the concentration of acetic acid formed, b. The hydrolysis of an ester such as ethyl acetate illustrates a himolecular reaction that gives sodium acetate and ethanol as the product from which second.order rate constant can be calculated CH3COOCH+NaOH−>CH COONa+CHOH present in it direct or indirect titration with a standard solution of an acid.A B Products 2.303 − logb(a−x) xloga(b−x) If 'b' moles of ester and 'a' moles of NaOH are taken initially, then a'a Vo (the volume of acid equivalent to the amount of NaOH present initially) Ca x) a Vr(the volume of acid equivalent to the amount of NaOH present at time, 1)