

$\text{H}_2\text{SO}_4 - \text{HS}^- - \text{H}_3\text{PO}_4 \text{ H}_2\text{S} - \text{HSO}_3^- - \text{HS}^- - \text{HS}^- -$
 $\text{H}_2\text{S} (1.008 / 1 \text{ mol}) = 2016 \text{ H}_2\text{S} (32.065 / 1 \text{ mol}) -$
 $32.065 \times 2.016 + 32.065 + 63.996 = 98.077 \text{ g/mol} = 63.996$
 $\text{H}_2\text{S} (1.008 / 11 \text{ mol}) = 2016 \text{ H}_2\text{S} (32.065 / 1 \text{ mol}) =$
 $32.065 \text{ H}_2\text{S} (15.999 / 1 \text{ mol}) = 47.997 \text{ g/mol}$