

Addition of strong acid or base to buffers When strong acid is added to a buffer solution containing HA and A⁻, the H⁺ consumes A⁻ to form HA. So, [HA] increases while [A⁻] decreases. Two steps are required to calculate the resulting pH – first, a stoichiometry problem for the neutralization; second, an equilibrium problem for the new concentrations. Here, [HA] decreases while [A⁻] increases.