

K_w at 25°C is 1.0×10^{-14} . Likewise, the weaker an acid is, the smaller its K_a and the stronger its conjugate base (with a larger K_b). The stronger an acid is, the larger its K_a and the weaker its conjugate base (with a smaller K_b).—Anions of Weak Acids Anion of weak acids react with water to some extent to form OH^- (and the conjugate acid). This relationship allows you to find K_b for the conjugate base of a weak acid or K_a for the conjugate acid of a weak base.