

Similarly the concentration of hydroxide ion,  $\text{OH}^-$ , of a solution is commonly expressed in term of the pOH of the solution , which is define as the negative logarithm of  $\text{OH}^-$   $\text{pOH} = -\log ( \text{OH}^- )$  The hydroxide ion concentration can be obtained from pOH of the solution using the equation :  $(\text{OH}^-) = 10^{-\text{pOH}}$  Additionally the pH and pOH of any aqueous solution are related as are the hydrogen and the hydroxide ion concentration , the relative equation are :  $(\text{H}^+)(\text{OH}^-) = 1 \times 10^{-14}$   $\text{pH} + \text{pOH} = 14$