Introduction In this practical work, we focus on the determination of the unknown concentration of oxalic acid ($H_2C_2O_4$) through a redox titration process. The principle underlying this experiment hinges on a redox reaction, where transfer of electrons occurs between oxidizing and reducing agents. Specifically, oxidation involves electron loss, while reduction pertains to electron gain. By employing potassium permanganate ($KMnO_4$) as the titrant, and using the titration method without an external indicator due to the self-indicating nature of $KMnO_4$, we aim to identify the precise point of reaction through a color change. This practical exercise is crucial for understanding redox processes and their applications in analytical chemistry.

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